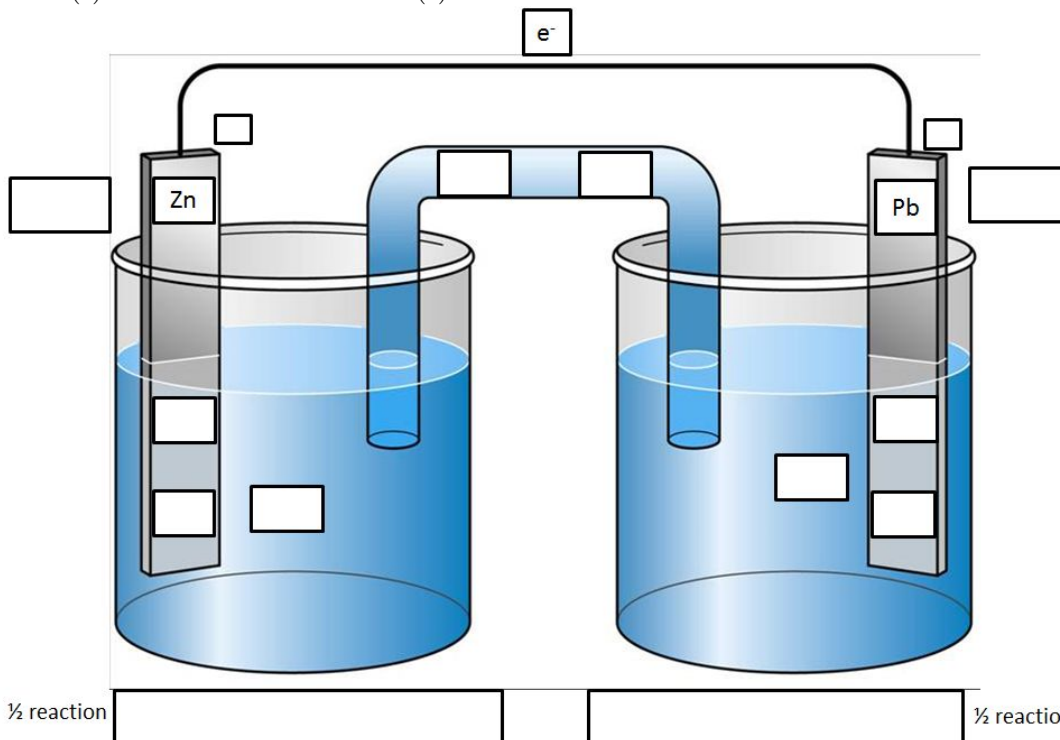


Name: _____

Date: _____

- [6 pt] 1. Using the cell drawn below. Label the following: electrodes (+/-), species in solution, anode, cathode, direction of the electron flow in the wire and the flow of all ions in solution. Assume that NaNO_3 is used in the salt bridge. Write the balance 1/2-reactions under each electrode. The galvanic cell reaction is $\text{Zn(s)} + \text{Pb}^{+2} \longrightarrow \text{Zn}^{+2} + \text{Pb(s)}$



- [6 pt] 2. Repeat the previous question for a cell whose reaction is: $3\text{Cu}^{+2} + 2\text{Cr(s)} \longrightarrow 3\text{Cu(s)} + 2\text{Cr}^{+3}$

