



[4 pt] 7. Define the term 'Electronegativity' and complete the following table.

Type of Bond	$\Delta EN$	Example
Ionic		
Polar Covalent		
Covalent		

[4 pt] 8. Explain the following two trends in Electronegativity.

(a) Electronegativity (Increases, Decreases or Same) across a row (Li vs F)? Explain.

(b) Electronegativity (Increases, Decreases or Same) down a column (Li vs Cs)? Explain.

[5 pt] 9. Classify the bond between the following pairs of elements as primarily (I)onic, (P)olar covalent, or (N)onpolar covalent in nature. Show work to support your answers.

(a) NaCl 9(a) \_\_\_\_\_

(b) CH<sub>4</sub> 9(b) \_\_\_\_\_

(c) CCl<sub>4</sub> 9(c) \_\_\_\_\_

(d) CaO 9(d) \_\_\_\_\_

(e) N<sub>2</sub> 9(e) \_\_\_\_\_