

Name: _____

Date: _____

[4 pt] 1. List the 7 metalloids. Why is important to know them?

[3 pt] 2. Define Compound. What is the driving force behind the formation of compounds?

[7 pt] 3. What are the differences between Ionic compounds and Covalent/Molecular compounds?

Property	Ionic	Covalent/Molecular
1. Formed by		
2. Between		
3. Bond Strength		
4. Melting Point		
5. Solubility		
6. Conductivity		
5. Structure		

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[5 pt] 4. Identify the following compounds as either (I)onic or (M)olecular. Explain.

(a) Potassium Chloride (KCl) 4(a) _____

(b) Sugar (C₁₂H₂₂O₁₂) 4(b) _____

(c) Nitrogen Dioxide (NO₂) 4(c) _____

(d) Iron (III) Bromide (FeBr₃) 4(d) _____

(e) Water (H₂O) 4(e) _____

[3 pt] 5. What are polyatomic ions? What is the name **AND** formula of the only positively charged polyatomic ion? Give an example of a negatively charged polyatomic ion (name and formula).

[18 pt] 6. Complete the following table, filling in each box with the proper formula.

	Br ⁻	O ⁻²	NO ₃ ⁻	PO ₄ ⁻³	CO ₃ ⁻²
K ⁺					
Mg ⁺²					
Al ⁺³					
Zn ⁺²					
H ⁺					