Complete the following reactions. Circle the most favored products.

2.
$$+ 2 \operatorname{Ag}^{+}_{(aq)} + \operatorname{NH}_{3(aq)} \xrightarrow{[0]}$$

$$6. \qquad \qquad + - \text{OH} \qquad \underbrace{ \left[\text{dry HCI} \right] }$$

7.
$$+ 2 Ag^{+}_{(aq)} + NH_{3(aq)}$$
 [0]

9.
$$+ 2 \text{ Ag}^+_{(aq)} + \text{NH}_{3(aq)} \xrightarrow{[0]}$$

10.
$$\uparrow$$
 + HCN [OH]

12.
$$+ 2 \text{ Ag}^+_{(aq)} + \text{NH}_{3(aq)} \xrightarrow{[0]}$$

14.
$$+ 2 Cu^{+2}_{(aq)} + NaOH$$

17.
$$\bigcirc$$
0 + \bigcirc 0 \bigcirc H⁺]

18.
$$\frac{\text{H}_2/\text{Ni}}{\Delta}$$

19.
$$\rightarrow$$
 $+$ \bigcirc $O_{2 (g)} \rightarrow$

$$20. \quad \stackrel{\text{OH}}{\longrightarrow} \quad \stackrel{\text{[o]}}{\longrightarrow} \quad$$

CHE102 - Extra Practice E2 - S18 - Ver. 1

Question 1:
$$+ H_2O$$
 Rate: Slow

Question 2: No Reaction

Question 4:
$$+ H_2O$$

Question 5:
$$\underline{\hspace{0.2cm}}$$
 co₂ (g) + $\underline{\hspace{0.2cm}}$ H₂O (g) + Heat Balance: $(1,4,3,3)$

Question 6:
$$/$$
 + H_2O

Question 7:
$$\longrightarrow$$
 \circ \circ \circ +2 Ag (s) Visible Change: clear \longrightarrow Silver Mirror

Question 8:
$$\stackrel{\mathsf{OH}}{\smile}$$
 + $\stackrel{\mathsf{OH}}{\smile}$ + $\stackrel{\mathsf{OH}}{\smile}$ + $\stackrel{\mathsf{OH}}{\smile}$ + $\stackrel{\mathsf{OH}}{\smile}$ + $\stackrel{\mathsf{OH}}{\smile}$

Question 9: —
$$\mathring{\text{onh}}_{4}$$
 +2 Ag (s) Visible Change: clear \longrightarrow Silver Mirror

Question 11:
$$+ H_2O$$

Question 12:
$$+2 \text{ Ag}_{(s)}$$
 Visible Change: clear \longrightarrow Silver Mirror

Question 14:
$$\leftarrow$$
 $+Cu_2O_{(s)}$ $+H_2O$ Visible Change: clear blue \longrightarrow brick-red

Question 16: $__{\text{CO}_2}$ (g) + $__{\text{H}_2\text{O}}$ (g) + Heat Balance: (1,4,3,3)

Question 19:
$$__{\text{CO}_2}$$
 (g) + $__{\text{H}_2\text{O}}$ (g) + Heat Balance: $(1,9,6,7)$

Question 20: No Reaction