

CHE 102 – Review – E1 – Reactions

1. Reaction Rules – General

<p>List 5 Types of Reactions:</p> <ol style="list-style-type: none"> 1. 2. 3. 4. 5. 	<p>List 5 Reaction Arrow + Meaning:</p> <ol style="list-style-type: none"> 1. 2. 3. 4. 5.
<p>List 2 Named Rules for reactions & write the rule & give an example illustrating the rule</p> <ol style="list-style-type: none"> 1. 2. 	
<p>Free Radicals:</p> <ol style="list-style-type: none"> 1. Definition: 2. Draw an example (for Cl_2): 3. Show how they are formed for $\text{C}=\text{C}$: 	<p>What is meant by the term “unique” hydrogens. What two “general” rules will help you identify them? Draw and example illustrating them.</p> <ol style="list-style-type: none"> 1. Definition: 2. Rule 1: 3. Rule 2:
<p>Draw a methyl, 1°, 2°, and 3° Free Radical. Order them from least stable to most stable:</p>	<p>Draw a methyl, 1°, 2°, and 3° Carbocation. Order them from least stable to most stable:</p>
<p>Named Reactions: Give an example reaction and the visual change that occurs (ie how you know a positive result occurred):</p> <ol style="list-style-type: none"> 1. Bromine Test: 2. Baeyer Test: 	

2. Reaction Types - Specific

Reaction Type	Reaction Arrow	Recognition	Multiple Products	Favored Product	Examples	Miscellaneous
Substitution (4 Types)						
Elimination (2 Types)						
Addition (4 Types)						
Combustion						
Oxidation						
Aromatics (2)						

