

Name: \_\_\_\_\_

Date: \_\_\_\_\_

[15 pt] 1. Name the following molecules:

(a)  $\text{Sb}_2(\text{S}_2\text{O}_3)_5$  1(a) \_\_\_\_\_(b)  $\text{Au}(\text{IO}_2)_2$  1(b) \_\_\_\_\_(c)  $\text{V}(\text{CN})_3$  1(c) \_\_\_\_\_(d)  $\text{Pb}_2\text{C}_2\text{O}_4$  1(d) \_\_\_\_\_(e)  $\text{Hg}(\text{PO}_4)_2$  1(e) \_\_\_\_\_(f)  $\text{Co}(\text{Cr}_2\text{O}_7)_2$  1(f) \_\_\_\_\_(g)  $\text{VClO}_3$  1(g) \_\_\_\_\_(h)  $\text{Sn}(\text{HSO}_3)_2$  1(h) \_\_\_\_\_(i)  $\text{Mn}(\text{CO}_3)_2$  1(i) \_\_\_\_\_(j)  $\text{CuF}_2$  1(j) \_\_\_\_\_(k)  $\text{Sn}(\text{C}_2\text{O}_4)_2$  1(k) \_\_\_\_\_(l)  $\text{Pb}(\text{MnO}_4)_5$  1(l) \_\_\_\_\_(m)  $\text{Au}_2\text{SO}_4$  1(m) \_\_\_\_\_(n)  $\text{V}_2(\text{CO}_3)_3$  1(n) \_\_\_\_\_(o)  $\text{Cr}_3(\text{PO}_4)_2$  1(o) \_\_\_\_\_

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[15 pt] 2. Write the formula for the following molecules:

(a) Vanadium (VI) Sulfate 2(a) \_\_\_\_\_

(b) Copper (II) Nitrite 2(b) \_\_\_\_\_

(c) Mercury (III) Carbonate 2(c) \_\_\_\_\_

(d) Arsenic (V) Chlorate 2(d) \_\_\_\_\_

(e) Chromium (II) Thiosulfate 2(e) \_\_\_\_\_

(f) Tin (I) Sulfate 2(f) \_\_\_\_\_

(g) Mercury (III) Hydroxide 2(g) \_\_\_\_\_

(h) Titanium (II) Hydrogen Phosphate 2(h) \_\_\_\_\_

(i) Manganese (IV) Oxide 2(i) \_\_\_\_\_

(j) Antimony (V) Phosphate 2(j) \_\_\_\_\_

(k) Cobalt (II) Acetate 2(k) \_\_\_\_\_

(l) Iron (III) Dichromate 2(l) \_\_\_\_\_

(m) Vanadium (II) Nitrate 2(m) \_\_\_\_\_

(n) Titanium (IV) Oxide 2(n) \_\_\_\_\_

(o) Lead (VI) Carbonate 2(o) \_\_\_\_\_