

Molecular Weight of all compounds:

1. $\text{H}_2(\text{g})$: 2.0158g

5. $\text{Br}_2(\text{l})$: 159.8g

9. $\text{PO}_4(\text{s})$: 94.97g

2. $\text{Sn}(\text{s})$: 118.71g

6. $\text{KBr}(\text{aq})$: 119g

10. $\text{NaCl}(\text{aq})$: 58.44g

3. $\text{H}_2\text{C}_2\text{O}_4(\text{aq})$: 90.0358g

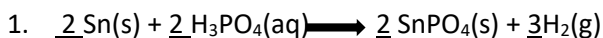
7. $\text{H}_3\text{PO}_4(\text{aq})$: 97.9937g

11. $\text{NaOH}(\text{aq})$: 39.9979g

4. $\text{MgC}_2\text{O}_4(\text{s})$: 112.33

8. $\text{SnPO}_4(\text{s})$: 213.68g

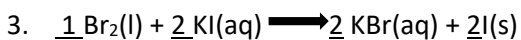
Complete the following equations- make sure to include the proper significant figures and units



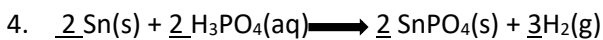
Provide the mol-mol ratio of the indicated molecules SnPO_4 to Sn _____



How many moles of MgC_2O_4 is reacted from 6.0 moles of H_3AsO_4 ? _____



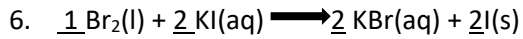
If you need to make 10.0 moles of KBr , how many moles of Br_2 are required? _____



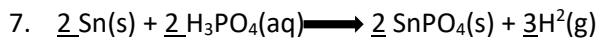
How many grams of Sn are produced when there are 10.5 grams of H_2 ? _____



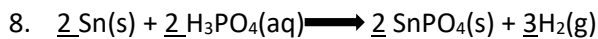
How many grams of MgC_2O_4 are consumed to produce 720.0 grams of $\text{H}_2\text{C}_2\text{O}_4$? _____



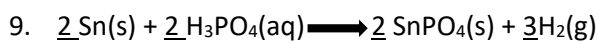
How many grams of KBr must be combusted to produce 45.0 kg of Br_2 ? _____



How many moles of H_2 are there if 85.0 grams are consumed? _____



If there are 46. grams of H_3PO_4 , how many moles are there? _____



If there are 4.0 moles of SnPO_4 , how many grams will there be? _____

10. 60.0 moles $\text{NaOH}_{(\text{aq})}$ to molecules

11. 160. Millimole $\text{H}_2\text{SO}_{4(\text{aq})}$ to molecules

12. 75.0 grams $\text{PO}_{4(\text{s})}$ to molecules

13. 120 kg $\text{NaCl}_{(\text{aq})}$ to molecules

14. How many atoms of Sn of SnPO_4 are in 7.7 grams of $\text{SnPO}_{4(\text{s})}$?

15. How many grams of 8.5 atoms of $\text{NaOH}_{(\text{aq})}$?

16. How many Oxygen atoms are in 6.7 moles of $\text{Fe}(\text{OH})_{3(\text{s})}$?

17. How many Oxygen atoms are in 8.02×10^{23} moles of $\text{H}_2\text{O}_{(\text{l})}$?
